

(19)



Europäisches Patentamt  
European Patent Office  
Office européen des brevets



(11) Publication number:

**0 520 419 A2**

(12)

## EUROPEAN PATENT APPLICATION

(21) Application number: **92110661.3**

(51) Int. Cl.<sup>5</sup>: **C07C 69/753, C07C 67/347,  
C07D 303/40**

(22) Date of filing: **25.06.92**

(30) Priority: **26.06.91 US 721799**

(43) Date of publication of application:  
**30.12.92 Bulletin 92/53**

(84) Designated Contracting States:  
**AT BE CH DE DK ES FR GB GR IT LI LU MC  
NL PT SE**

(71) Applicant: **UNION CARBIDE CHEMICALS &  
PLASTICS TECHNOLOGY CORPORATION**  
39 Old Ridgebury Road  
Danbury, Connecticut 06817-0001(US)

(72) Inventor: **Koleske, Joseph Victor**  
1513 Brentwood Road  
Charleston 25314, West Virginia(US)  
Inventor: **Argyropoulos, John Nicholas**  
35 Michael Street  
Scott Depot 25560, West Virginia(US)  
Inventor: **Smith, Oliver Wendell**  
850 Walters Road  
Charleston 25314, West Virginia(US)

(74) Representative: **von Hellfeld, Axel, Dr.**  
**Dipl.-Phys.**  
**Wuesthoff & Wuesthoff Patent- und**  
**Rechtsanwälte Schweigerstrasse 2**  
**W-8000 München 90(DE)**

(54) **Production of unsaturated cycloaliphatic esters and derivatives thereof.**

(57) Unsaturated cycloaliphatic esters like higher hydrocarbyl, functionally substituted or polyunsaturated cyclohex-3-ene carboxylates, made directly by cycloaddition of dienes with dienophilic (meth/eth)acrylates, and their derivatives like epoxides and urethanes, provide useful thermal and radiation curable coatings, inks, sealants, adhesives, solvents, acid scavengers, and intermediates for other uses.

EP 0 520 419 A2

## Brief Summary of the Invention

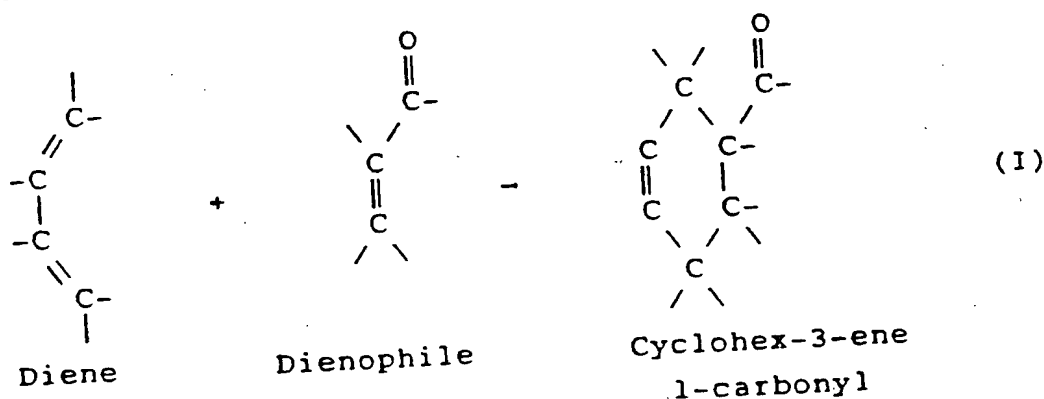
## Technical Field

This invention relates to unsaturated cycloaliphatic esters produced by cycloaddition reactions and derivatives thereof, and more particularly to higher hydrocarbyl, functionally substituted or poly-unsaturated cyclohex-3-ene carboxylates made by reacting dienes and dienophilic (meth/eth)acrylates, and especially 3,4-epoxy derivatives thereof useful in curable coatings.

## Background of the Invention

Cycloaddition reactions of dienes, i.e., compounds having conjugated carbon-carbon double bonds, such as butadiene and cyclopentadiene, with dienophiles, i.e.,  $\alpha,\beta$ -unsaturated carbonyl compounds, such as acrylic acid, produce unsaturated cycloaliphatic carbonyl compounds. This reaction, also called the

Diels-Alder reaction, is shown in Equation I.



Such reactions are well known and described in the literature. For example, U.S. Patent No. 1,944,732 (Diels et al.) discloses reacting dienes with acrolein, acrylic acid and ethylidene ketone compounds. Dienophilic esters, such as ethyl acrylate, are disclosed by Roberts et al., in *Basic Principles of Organic Chemistry*, Pages 262-265, W. A. Benjamin, Inc., New York (1964). Additional studies of reactions between dienes and dienophiles are disclosed by Kojima et al., in "Aluminum Chloride Catalyzed Diene Condensation. V. Selectivity Reactivity Relationship of Dienophiles toward Butadiene, Isoprene and 2-Trifluoromethylbutadiene", *Journal of Organic Chemistry*, 35, Pages 1342-1348 (1970). Reactions of 1-phenylseleno-2-trimethylsilyloxy-4-methoxy-1,3-butadiene and dienophiles are disclosed by Danishefsky et al., in "A Diels-Alder Route to Functionalized Cyclohexadienones", *Journal of Organic Chemistry*, Volume 42, Pages 1819-1821 (1977). Reaction of 2-chloro-3-phenyl-1,3-butadiene with the dienophiles of (meth)acrylic and cinnamic acids, or their methyl, ethyl and butyl esters, are disclosed by Martirosyan et al., in *Arm. Khim. Zh.*, 24(8), Page 697 (1971), *Chemical Abstracts*, 76:3514a.

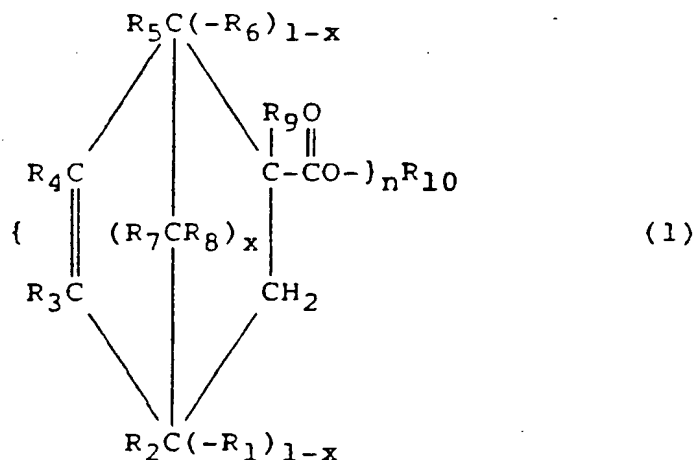
Other cycloaddition reactions and derivitizations of cycloaliphatic compounds are also described in the literature. U.S. Patent No. 2,794,812 (Phillips et al.) discloses reacting dienes with dienophiles which are  $\alpha,\beta$ -unsaturated aliphatic acids followed by esterification with alcohols and oxidation to produce 7-oxabicycloheptane 3-carboxylic acids and esters. U.S. Patent No. 2,716,123 (Frostick et al.) discloses cyclic unsaturated aldehydes, made by reacting diene with  $\alpha,\beta$ -unsaturated aliphatic aldehydes. These cyclic aldehydes are coupled by the Tischenko reaction between aldehydes to form the unsaturated dicycloaliphatic esters, which are epoxidized to make diepoxycycloaliphatic esters. U.S. Patent No. 2,745,847 (Phillips et al.) discloses oxidizing unsaturated cycloaliphatic aldehydes to the corresponding unsaturated cycloaliphatic carboxylic acid, which are then reacted with glycols of aliphatic or oxyalkylene groups, i.e., dihydric alcohols to form unsaturated dicycloaliphatic esters which are then epoxidized to the corresponding diepoxides. Similarly, U.S. Patent No. 2,750,395 (Phillips et al.) discloses reducing unsaturated cycloaliphatic aldehydes to the corresponding unsaturated cycloaliphatic alcohols, which are reacted with dicarboxylic acids to form unsaturated dicycloaliphatic esters which are epoxidized to the corresponding diepoxides. Various cyclohexene carboxylic esters, made by reacting cyclohexene carboxylic acid with various alkyl, phenyl and phenylene ester chlorides are disclosed by Pishnamazzade et al., in *Zh. Org.*

Khim., 9(4), Pages 715-719 (1973), Chemical Abstracts, 79:31748k. These compounds are then epoxidized to form the corresponding epoxide. Bis(cyclohexene carboxylates), made by reacting cyclohexene carboxylic acid with ester dichlorides, and then oxidizing or halogenating to epoxy or dihalo products, are disclosed by Pishnamazade et al., in Zh. Org. Khim, 10(4), Pages 712-717 (1974), Chemical Abstracts, 81:25473s; and Kerimov et al., in Azerb. Khim. Zh., (3), Pages 52-57 (1983), Chemical Abstracts, 100:191410z.

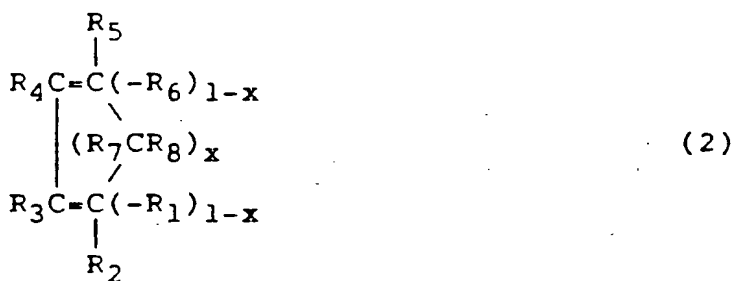
Production of various, such as functionally substituted or multifunctional, unsaturated cycloaliphatic compounds has involved multiple reactions. For example, diene is reacted with dienophilic acid or aldehyde to produce unsaturated cycloaliphatic acid or aldehyde. The aldehyde is then reduced to the corresponding alcohols, oxidized to the corresponding acid, or coreacted to produce corresponding di(esters). Through further reactions, the unsaturated cycloaliphatic acids are esterified either directly with alcohols or using base to form the corresponding salt and reaction with halides for further derivatization. These procedures require multiple reactions and/or produce by-products, such as salts, requiring disposal. Of course, a by-product, KCl, which had to be removed and properly disposed of in additional steps, was formed with the earlier-cited reaction sequence of Kerimov et al. It would be desirable if a more direct procedure could be provided for producing certain unsaturated cycloaliphatic esters.

### Summary of the Invention

This invention relates to a process for producing unsaturated cycloaliphatic esters represented by Formula 1.



In Formula 1:  $\underline{n}$  is at least 1;  $R_1$ - $R_8$  are each, independently, hydrogen,  $C_1$ - $C_{10}$  hydrocarbyl with or without halo substitution, halo, cyano, or silyl;  $R_9$  is hydrogen, methyl, or ethyl;  $R_{10}$  is hydrocarbyl or oxyhydrocarbyl, provided  $R_{10}$  has at least 5 carbon atoms or is oxyhydrocarbyl or is  $-CH=CH_2$  when  $\underline{n}$  is 1; and  $\underline{x}$  is 0 or 1. The process comprises a reaction using diene represented by Formula 2.



In Formula 2,  $R_1$ - $R_8$  and  $\underline{x}$  are as defined in Formula 1. The diene is reacted with a dienophilic (meth/eth)-acrylate represented by Formula 3.



In Formula 3,  $n$  and  $\text{R}_9$ - $\text{R}_{10}$  are as defined in Formula 1.

This invention also relates to processes for epoxidizing, hydrogenating and transesterifying unsaturated cycloaliphatic esters represented by Formula 1 above. The unsaturated cycloaliphatic esters can be prepared by reacting a compound with a system of conjugated carbon-carbon multiple bonds with an acrylate through a Diels-Alder reaction as indicated above. Then, in subsequent reactions, these unsaturated cycloaliphatic esters can be either oxidized with an oxidizing agent such as peracetic acid to a cycloaliphatic epoxide or reduced to the saturated compound by means of hydrogenation. The compounds of the invention can also be produced by transesterifying the unsaturated or saturated compound or the epoxidized compound with other suitable compounds. The compounds of this invention can be made by a variety of processes that include Diels-Alder reactions followed by subsequent epoxidation or hydrogenation, Diels-Alder reactions followed by epoxidation or hydrogenation and subsequent transesterification or transesterification, Diels-Alder reactions followed by transesterification or transesterification and subsequent epoxidation or hydrogenation and the like.

This invention further relates to unsaturated cycloaliphatic esters of Formula 1 above and derivatives thereof including derivatives formed by epoxidation, hydrogenation, transesterification and the like.

#### Detailed Description

It has been found that unsaturated cycloaliphatic esters having higher hydrocarbyl or functionally substituted alcohol moiety or polyunsaturation can be made from diene with dienophilic (meth/eth)acrylate ester. The esters are produced without the need for multiple reaction steps or unnecessary by-product. Thus, in a single reaction step, a compound is made, without any by-product formation, that previously required up to three reaction steps for its manufacture, i.e., first formation of the cyclohex-3-ene carboxylate, then reaction of the carboxylate with a base to form the potassium salt, and finally reaction of the potassium salt with an alkyl or other chloride (see Kerimov et al., noted previously).

The term "hydrocarbyl" is used in this specification to mean any radical containing hydrogen and carbon atoms. The term "oxyhydrocarbyl" is used in this specification to mean any radical containing oxygen, hydrogen and carbon atoms.

#### Cycloaddition

Dienes which may be used include those shown in Formula 2 wherein  $\text{R}_1$ - $\text{R}_8$  and  $x$  are as defined in Formula 1. When  $x$  is 0, the diene is an acyclic, conjugated diene, represented by Formula 4.



In Formula 4,  $\text{R}_1$ - $\text{R}_6$  are as defined in Formula 1. When  $x$  is 1, the diene is a cyclopentadiene, represented by Formula 5.



In Formula 5,  $\text{R}_2$ - $\text{R}_5$  and  $\text{R}_7$ - $\text{R}_8$  are as defined in Formula 1. In Formulas 2, 4 and 5, four or more of  $\text{R}_1$ - $\text{R}_8$  are preferably hydrogen, more preferably  $\text{R}_3$  or  $\text{R}_4$  are hydrogen, and most preferably  $\text{R}_3$  and  $\text{R}_4$  are both hydrogen.

Illustrative diene substituents, represented by  $\text{R}_1$ - $\text{R}_8$  in Formulas 2, 4 and 5, include, among others: hydrogen; alkyl such as methyl, ethyl, propyl, butyl, etc.; aryl such as phenyl; cycloalkyl such as cyclohexyl; halo-substituted hydrocarbyl such as polyfluoromethyl; halo such as fluoro, chloro, bromo, iodo, etc.; cyano; and silyl such as polymethylsilane.

Suitable dienes include, among others, one or more: 1,3-butadiene; homologs of butadiene; isoprene; homologs of isoprene; 1,3-pentadiene; cyclopentadiene; myrcene; phellandrene; 1,3-hexadiene; 2,4-hexadiene; 1,3,5-hexatriene; 1,3-octadiene; 2,4-octadiene; 3,5-octadiene; 1,3,5,7-octatetraene; 2-trifluoromethyl-1,3-butadiene; 1-methyl-1,3-butadiene; 2-methyl-1,3-butadiene; 1-phenyl-1,3-butadiene; 2-phenyl-1,3-butadiene; 1-cyclohexyl-1,3-butadiene; 2-cyclohexyl-1,3-butadiene; 1-cyclohexyl-1,3-isoprene; 2-cyclohexyl-1,3-isoprene; 1-chloro-1,3-butadiene; 2-chloro-1,3-butadiene; 1-cyano-1,3-butadiene; 2-cyano-1,3-butadiene; 2,3-dimethyl-1,3-butadiene; and the like. Preferred dienes include 1,3-butadiene, isoprene and cyclopentadiene.

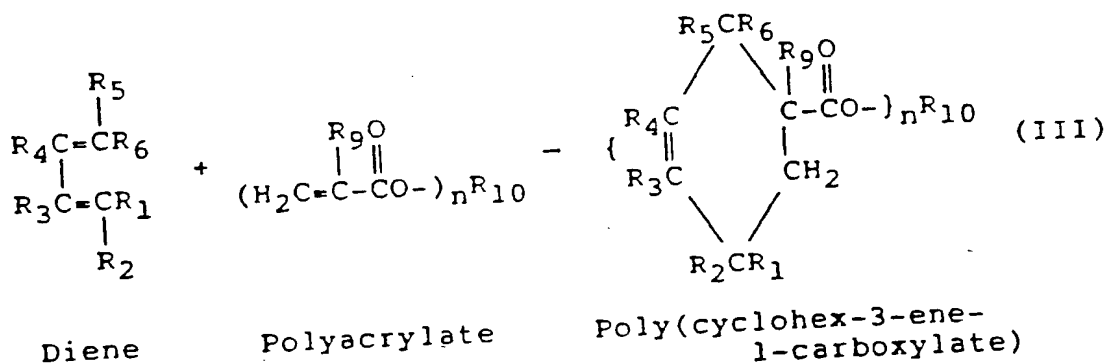
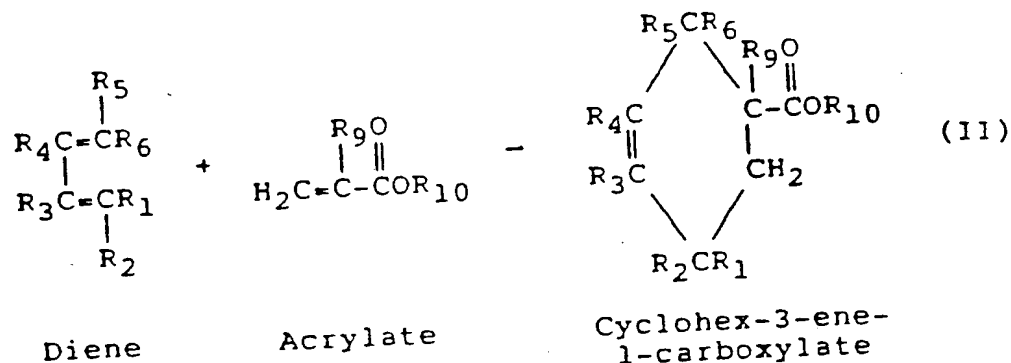
Dienophilic (meth/eth)acrylates, i.e. dienophiles, are as shown in Formula 3. In Formula 3,  $\text{R}_9$ - $\text{R}_{10}$  and  $n$  are as defined in Formula 1. The number of acrylic groups, defined by  $n$ , is typically from 1 to about 10, and when polyacrylic is preferably from 2 to about 6.  $\text{R}_9$  is preferably hydrogen or methyl, and most preferably hydrogen.

Illustrative  $\text{R}_{10}$  groups include, among others:  $\text{C}_5$ + alkyl such as pentyl, hexyl, octyl, dodecyl, hexadecyl, and so on; aryl such as phenyl; cycloalkyl such as cyclohexyl; vinyl; alkoxy such as ethoxy; aryloxy such as phenoxy; hydroxyalkyl such as hydroxy ethyl, hydroxypropyl, 2-ethyl-2-hydroxyethyl, and so on; alkoxyated hydroxyalkyls, including addition products of hydroxyalkyls with alkylene oxides like ethylene oxide, propylene oxide, tetrahydrofurans, methoxytetrahydrofurans, lactones like caprolactone and valerolactone and propiolactone, such as  $-\text{CH}_2\text{CH}_2-\text{O}-(\text{CH}_2\text{CH}_2\text{O})_m-\text{H}$ ,  $-\text{CH}_2\text{CHCH}_3-\text{O}-(\text{CH}_2\text{CH}_2\text{O})_m-\text{H}$ ,  $-\text{CH}_2\text{CH}_2-\text{O}-(\text{CH}_2\text{CHCH}_3\text{O})_m-\text{H}$ ,  $-\text{CH}_2\text{CHCH}_3-\text{O}-(\text{CH}_2\text{CHCH}_3\text{O})_m-\text{H}$ ,  $-\text{CH}_2\text{CH}_2-\text{O}-((\text{CH}_2)_4\text{O})_m-\text{H}$ ,  $-\text{CH}_2\text{CHCH}_3-\text{O}-((\text{CH}_2)_4\text{O})_m-\text{H}$ ,  $-\text{CH}_2\text{CH}_2-\text{O}-((\text{CH}_2)_5\text{O})_m-\text{H}$ ,  $-\text{CH}_2\text{CHCH}_3-\text{O}-((\text{CH}_2)_5\text{O})_m-\text{H}$ ,  $-\text{CH}_2\text{CH}_2-\text{O}-(\text{CO}(\text{CH}_2)_5\text{O})_m-\text{H}$ ,  $-\text{CH}_2\text{CHCH}_3-\text{O}-(\text{CO}(\text{CH}_2)_5\text{O})_m-\text{H}$ , and the like where  $m$  is from 1 to about 25, preferably from 1 to about 10; or the residue formed by reacting (meth/eth)acrylic acid with polyhydroxy compound, i.e., compounds having 2 or more hydroxyl groups, providing polyvalent residue  $\text{R}_{10}$ . For example,  $\text{R}_{10}$  is  $\text{CH}_3\text{CH}_2\text{C}(\text{CH}_2\text{O})_3$  with  $n$  equal to 3 or is  $-\text{O}-\text{CH}_2\text{CH}_2-\text{O}-\text{CH}_2\text{CH}_2-\text{O}-$  with  $n$  equal to 2 when the polyhydroxy compound is trimethylolpropane or diethylene glycol, respectively. Illustrative polyhydroxy compounds include, among others, one or more: ethylene glycol; diethylene glycol; triethylene glycol; tetraethylene glycol; trimethylolpropane; pentaerythritol; 1,3-butylene glycol; polyethylene glycol; caprolactone; tripropylene glycol; poly(propylene glycol); Bisphenol A; linseed oil; soybean oil; as well as alkoxyated derivatives of such polyhydroxy compounds.

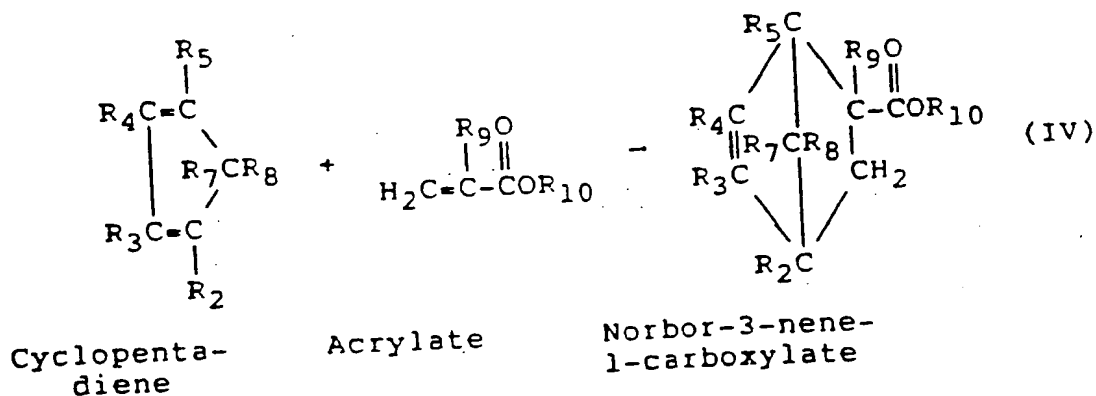
Illustrative dienophilic (meth/eth)acrylates, by which is meant acrylates, methacrylates, and ethacrylates, include, among others, one or more: esters of (meth/eth)acrylic acid with monohydric and polyhydric compounds, such as pentyls, hexyls, octyls, decyls, and the like; neopentyl diacrylate; esterdiol diacrylates such as 2,2-dimethyl-3-hydroxypropyl-2,2-dimethyl-3-hydroxypropionate diacrylate; trimethylolpropane triacrylate; pentaerythritol di-, tri-, and tetraacrylate; hydroxyethyl acrylate; hydroxypropyl acrylate; (poly)-caprolactone acrylates; ethoxyated acrylates; propoxyated acrylates; glycerol acrylates; triethylene glycol diacrylate; tetraethylene glycol diacrylate; ethoxyethyl acrylate; cyclohexyl acrylate; 2-phenoxyethyl acrylate; isobornyl acrylate; 1,3-butylene glycol diacrylate; 1,4-butanediol diacrylate; 1,6-hexanediol diacrylate, glycidyl acrylate; dipentaerythritol acrylates; poly(ethylene glycol) acrylates; caprolactone di-, tri-, and tetraacrylates, as well as other lactone acrylates; tripropylene glycol diacrylate; poly(propylene glycol) acrylates; ethoxyated or propoxyated Bisphenol-A diacrylates; alkoxyated esterdiol diacrylates such as ethoxyated or propoxyated 2,2-dimethyl-3-hydroxypropyl-2,2-dimethyl-3-hydroxypropionate diacrylates; ac-

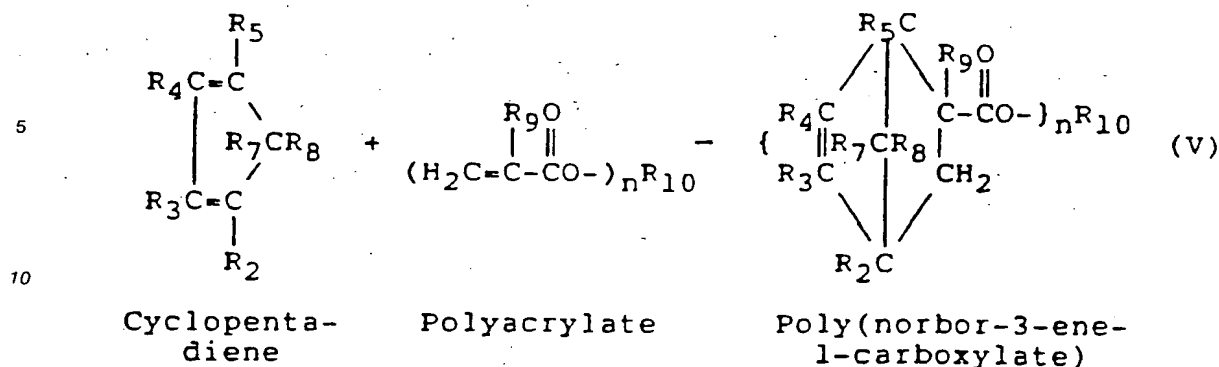
rylates of caprolactone reacted with esterdiols; ethoxylated or propoxylated trimethylolpropane triacrylate; ethoxylated or propoxylated pentaerythritol di-, tri-, or tetraacrylate; unsaturated polyesters containing ethylenic unsaturation from maleic, fumaric, citraconic acids and the like; unsaturated dicarboxylic acids; urethane acrylates of various types; epoxy acrylates; acrylated polybutadiene; acrylated linseed oil; acrylated soybean oil; and the like. In one embodiment, mixtures of multifunctional and monofunctional dienophiles can be used with monofunctional dienophile having  $R_{10}$  hydrocarbonyl groups with 1 or more carbon atoms.

The diene reacts with the dienophilic (meth/eth)acrylate, which undergoes 1,4-cycloaddition, to produce cyclohex-3-ene-1-carboxylates. Of course, as is understood by those skilled in the art, a variety of isomers can be obtained. Acyclic dienes, when  $x$  is 0, react with monofunctional, when  $n$  is 1, or polyfunctional, when  $n$  is more than 1, (meth/eth)acrylates as shown in Equation II and III, respectively.



Cyclic dienes, when  $x$  is 1, react with monofunctional or polyfunctional (meth/eth)acrylates as shown in Equations IV and V, respectively.





The relative proportion of diene to dienophile is not narrowly critical and may vary depending on the degree of functionality, i.e. number of acrylic groups, in the dienophilic (meth/eth)acrylate. Usually, the amount of diene to dienophile (meth/eth)acrylate is from about 5:1 to about 0.1:1, preferably from about 3:1 to about 0.2:1 and most preferably from about 2:1 to about 0.5:1, moles of diene per mole of ethylenic unsaturation in the dienophilic (meth/eth)acrylate.

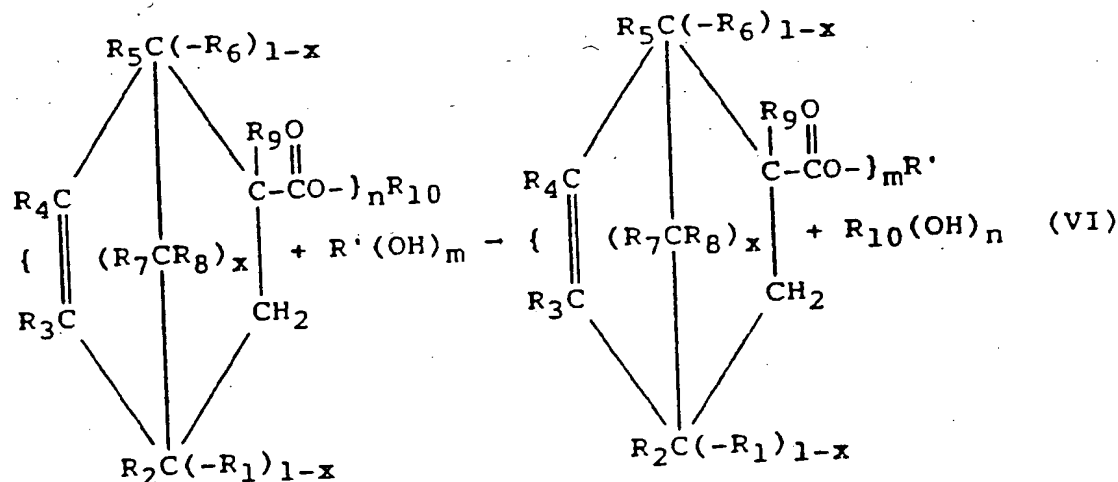
The particular reaction conditions for cycloaddition are not critical but can be any effective, including known, procedures for reacting dienes with dienophiles sufficient to produce the unsaturated cycloaliphatic esters. The reaction is usually carried out at temperature from about -100°C. to about 80°C., preferably -10°C. to about 50°C. The reaction may be conducted under atmospheric, i.e. ambient, subatmospheric or superatmospheric pressures, preferably at subatmospheric or atmospheric pressures, and most preferably at subatmospheric pressures.

Other ingredients may optionally be provided to the cycloaddition reaction. Catalyst may be used to improve reaction rate and the ratio of isomers produced, such as described by Sauer et al. in *Tetrahedron Letters*, (7), Pages 731-736 (1966). Suitable catalysts include, among others, one or more: aluminum chloride, aluminum (tri)chloride (di)etherate, boron trifluoride etherate, tin (IV)chloride, titanium tetrachloride, and others, and particularly methyl acrylate-aluminum chloride complex as described by Kojima et al. in the *Journal of Organic Chemistry*, 35(5), Page 1342 (1970). Suitable solvents may optionally be used, including, among others, one or more: methylene chloride, dioxane, methanol, ethanol, propanol, isopropanol, chloroform, dioxane, triethylamine, 1,2-dimethoxyethane, acetone, toluene, and others. Special solvents, such as lithium perchlorate in diethyl ether as described by Gaul in the *Journal of the American Chemical Society*, 112, Page 4595 (1990) may be used to accelerate the reaction.

The unsaturated cycloaliphatic esters so produced have mono- or polyunsaturation, as defined by  $m$  equal to 1 or greater than 1, respectively. When monofunctional, the esters have an alcohol moiety, of  $\text{R}_{10}$ , which is either hydrocarbyl having at least 5 carbon atoms or oxyhydrocarbyl, including functional groups like hydroxyl.

#### Transesterification

A particularly useful technique for preparing a variety of unsaturated cycloaliphatic carboxylates and particularly multifunctional unsaturated cycloaliphatic carboxylates, is transesterification. Unsaturated cycloaliphatic esters can undergo transesterification by reaction with hydroxyl-containing compound, as set forth in Equation VI.



In Equation VI,  $n$ ,  $\text{R}_1$ – $\text{R}_{10}$  and  $x$  are as defined in Formula 1;  $m$  identifying the number of hydroxyl groups in hydroxyl-containing compound  $\text{R}'(\text{OH})_m$ , is at least one, preferably from 1 to about 10; and  $\text{R}'$  is the residue of the hydroxyl-containing compound. Suitable hydroxyl-containing compounds include, among others, one or more: alcohols, such as methanol, ethanol, propanols, butanols, pentanols, hexanols, cyclohexanols, phenols, decanols, dodecanols, hexadecanols and others; glycols, such as ethylene glycol di-, tri-, tetraethylene glycols and other poly(ethylene glycols), propylene glycol, di-, tri-, and tetrapropylene glycol as well as other poly(propylene glycols); polyols such as trifunctional poly(propylene oxide)s including ethylene oxide-capped and  $\epsilon$ -caprolactone-capped propylene oxide polyols that contain up to about 25% by weight ethylene oxide or  $\epsilon$ -caprolactone oxide, random, block, and graft ethylene oxide/propylene oxide copolymers, polylactone polyols including  $\epsilon$ -caprolactone, various methyl caprolactone,  $\epsilon$ -valerolactone and methyl valerolactone, propiolactone polyols, poly(teramethylene oxide) polyols, polyester polyols including hexanediol adipates, butylene adipates, ethylene adipates, butylene succinates, polycarbonate polyols, hydroxyl-terminated polyethylenes, poly(vinyl alcohol), vinyl acetate/vinyl alcohol copolymers, styrene/allyl alcohol copolymers, hydroxyethyl cellulose, hydroxypropyl cellulose, cellulose acetates, hydroxylalkyl acrylates like hydroxyethyl acrylate and hydroxypropyl acrylate alone or after reaction with various amounts of propylene oxide or ethylene oxide, lactones like caprolactone acrylates that contain one or more free hydroxyl groups, and others.

The particular reaction conditions for transesterification are not critical but can be any effective, including known, procedures for reacting esters with hydroxy compounds sufficient to produce the unsaturated cycloaliphatic transesters. The reaction temperature may vary depending on the particular reactants used and is usually carried out at temperature from about 40°C. to about 30°C., preferably 60°C. to about 250°C. The reaction may be conducted under atmospheric, i.e. ambient, subatmospheric or superatmospheric pressures, preferably at subatmospheric or atmospheric pressures.

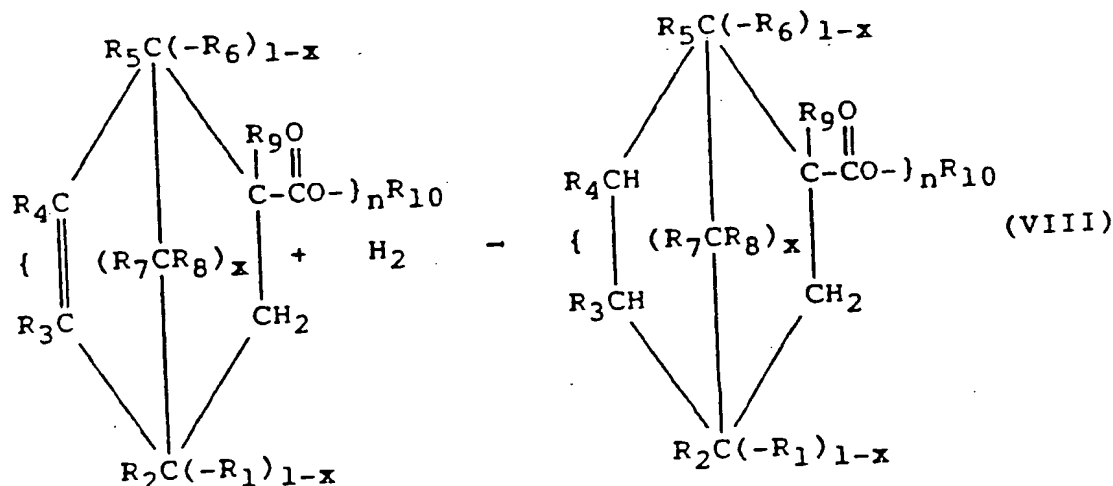
Other ingredients may optionally be provided to the transesterification reaction. Catalyst may be used. Suitable transesterification catalysts include, among others, one or more: tetrabutyltitanate; zinc acetate; titanium dioxide; sodium and potassium alkoxides; sodium and potassium phenoxides; lead oxide; ion exchange resins; and others. Time for conducting the transesterification reaction can vary generally from about 10 minutes to 40 or more hours, depending on the temperature employed and particular ingredients involved. Specific reaction conditions for transesterification may be determined based on known techniques, such as described by Tucek et al. in *Acta Polymerica*, 31(7), page 429 (1980).

#### Epoxidation

The unsaturated cycloaliphatic esters can be epoxidized to form cycloaliphatic epoxides and other, useful products, as shown in Equation VII.



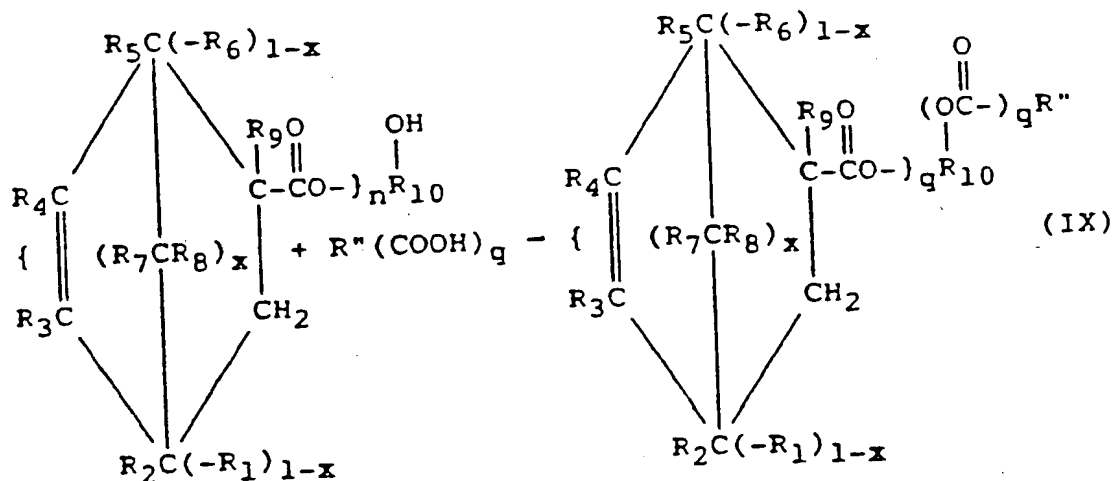




The particular reaction conditions for hydrogenation are not critical but can be any effective, including known, hydrogenation procedures sufficient to produce the saturated cycloaliphatic esters. The reaction may be carried out at generally temperature from about 10°C. to about 250°C., preferably from about 20°C. to about 200°C. The reaction may be conducted at pressures of about 0 to about 500 psig, and preferably from about 0 to about 200 psig.

Other ingredients may optionally be provided to the hydrogenation reaction. Suitable hydrogenation catalyst may be used including, among others: platinum; nickel; palladium; iron; cobalt molybdate on alumina; copper chromite; barium promoted copper chromite; tin-copper couple; zinc-copper couple; Raney-type compounds, such as Raney-nickel, aluminum-cobalt, aluminum-copper, and aluminum-nickel; and others.

Various unsaturated cycloaliphatic esters, such as those with hydroxyl groups, can be condensed with various mono- and polyfunctional carboxylic acids or anhydrides or their mixtures as set forth in Equation IX.



In Equation IX:  $n$ ,  $\text{R}_1$ - $10$  and  $x$  are as defined in Formula I including wherein  $\text{R}_{10}$  also has a hydroxyl group;  $\text{R}''$  is hydrocarbyl; and  $q$  is at least 1.  $\text{R}''(\text{COOH})_q$  is a monofunctional, when  $q$  is 1, or a polyfunctional, when  $q$  is more than 1, carboxylic acid, or may be the corresponding anhydride. Unsaturated cycloaliphatic esters having one or more free hydroxyl groups can be reacted with compounds like lactones such as  $\epsilon$ -caprolactone, methyl caprolactones,  $\epsilon$ -valerolactone, methyl valerolactones,  $\zeta$ -enantholactones,  $\beta$ -propiocaprolactone, and others, or with propylene oxide, ethylene oxide, tetrahydrofuran, or mixtures of these compounds to form new adducts, which can react with the carboxylic acids and/or anhydrides, with or without subsequent epoxidation.

Other derivatization reactions involving the unsaturated cycloaliphatic esters include transesterification reactions. The particular reaction conditions for transesterification are not critical and include conventional procedures known in the art.

It is understood that the various derivatization reactions can be conducted in any permissible sequence.

This invention is not to be construed as being limited to particular derivatization reactions or derivatization reaction sequences following the Diels-Alder reaction.

### Uses

The unsaturated cycloaliphatic esters, and derivatives thereof, especially epoxides, can be used alone or in combination with a variety of other ingredients such as hydroxyl-containing compounds, other cycloaliphatic epoxides and glycidyl epoxides, acrylates, and other suitable ingredients to form coatings, inks, adhesives, sealants, in the photoresist market area for production of printing plates, printed circuit boards, and similar products. The ingredients may be reacted together by actinic radiation, such as ultraviolet light, with suitable photoinitiators, under suitable, such as ambient, conditions, or by means of thermal energy, when suitable blocked or free initiators or catalysts are employed.

The cycloaliphatic epoxides of this invention are useful alone or in combination with other compounds to form a variety of articles of commerce including molded parts; coatings, inks, adhesives, and sealants cured by both thermal and ultraviolet light means as indicated above; acid scavengers; pharmaceutical products or intermediates for pharmaceutical products; flavors and fragrances; solvents; as well as other end uses. The epoxidized compounds of this invention may be used alone or in combination with other ingredients such as polyether, polyester, polycarbonate, and polylactone polyols; linear and cyclic vinyl ethers; anhydrides, and other ingredients known to those skilled in the art of product formulation. The unsaturated cycloaliphatic reaction products that are reduced or hydrogenated according to this invention have utility as solvents for various purposes such as dissolving polymers, decreasing viscosity of paint systems, paint removal, and the like.

Hydroxyl-containing compounds that can be used in combination with the epoxidized cycloaliphatic esters or other derivatives include, among others: alcohols such as butanols, pentanols, hexanols, decanols, ethoxy and propoxy alcohols such as ethoxyethanol, propoxyethanol, ethoxypropanol, propoxypropanol, ethoxybutanol, propoxybutanol, ethoxyethoxy- and propoxypropoxy-ethanol, propanol, or butanol; glycols such as ethylene glycol, di-, tri-, tetra-ethylene glycols and other poly(ethylene glycols), propylene glycol, di-, tri-, and tetrapropylene glycol as well as other poly(propylene glycols); polyols such as trifunctional poly(propylene oxide) polyols including ethylene oxide-capped and  $\epsilon$ -caprolactone-capped propylene oxide polyols that contain up to above 25% ethylene oxide, or caprolactone as the capping compound, random, block, and graft ethylene oxide/propylene oxide copolymers, polylactone polyols including  $\epsilon$ -caprolactone, various methyl caprolactone,  $\delta$ -valerolactone and methyl valerolactone, propiolactone polyols, poly-(tetramethylene oxide) polyols, polyester polyols including hexanediol adipates, butylene adipates, ethylene adipates, butylene succinates, polycarbonate polyols, hydroxyl-terminated polyethylenes, styrene/allyl alcohol copolymers, hydroxyethyl cellulose, hydroxypropyl cellulose; and others. In the radiation and thermal cure formulations, these compounds serve to increase reaction rate and, when molecular weight is sufficiently high, to toughen and/or flexibilize the formed product.

Suitable epoxides that can be used in combination with the epoxidized cycloaliphatic esters and if desired other ingredients including cycloaliphatic epoxides having an average of one or more epoxide groups per molecule, including, among others: 3,4-epoxycyclohexylmethyl 3,4-epoxycyclohexane carboxylates such as 3,4-epoxy-cyclohexylmethyl 3,4-epoxycyclohexane carboxylate, 3,4-epoxy-1-methylcyclohexylmethyl 3,4-epoxy-1-methylcyclohexane carboxylate, 6-methyl 3,4-epoxycyclohexylmethyl 6-methyl-3,4-epoxycyclohexane carboxylate, 3,4-epoxy-3-methyl-cyclohexylmethyl 3,4-epoxy-3-methylcyclohexane carboxylate; 3,4-epoxy-5-methylcyclohexylmethyl 3,4-epoxy-5-methyl-cyclohexane carboxylate, and as described in, for example, U.S. Patent No. 2,890,194; diepoxides of cycloaliphatic esters of dicarboxylic acids such as bis(3,4-epoxycyclohexylmethyl)oxylate, bis(3,4-epoxycyclohexylmethyl)adipate, bis(3,4-epoxy-6-methylcyclohexylmethyl)adipate, bis(3,4-epoxycyclohexylmethyl)pimelate, and as described in, for example, U. S. Patent No. 2,750,395; other cycloaliphatic diepoxides including 2-(3,4-epoxycyclohexyl-5,5-spiro-3,4-epoxy)cyclohexane-m-dioxane, halogen or monovalent hydrocarbon variations of this compound, and the like as further defined in U.S. Patent No. 3,318,822; cyclopentadiene diepoxide, cyclohexane diepoxide; and preferably 3,4-epoxycyclohexylmethyl-3,4-epoxycyclohexane carboxylate, bis(3,4-epoxycyclohexylmethyl)-adipate, 2-(3,4-epoxycyclohexyl-5,5-spiro-3,4-epoxy)cyclohexane-m-dioxane, or mixtures thereof.

Cycloaliphatic monoepoxides may also be used, in combination with the epoxidized cycloaliphatic esters, which may be an unsubstituted monoepoxide, such as cyclohexene oxide, or a monoepoxide

substituted with, for example, alkyl groups of 1 to 6 carbon atoms, halogen, ester groups, vinyl groups, and others. Suitable substituted monoepoxides include, among others: limonene monoepoxide; 4-vinyl cyclohexene monoepoxide; norbornene monoepoxide; alpha-pinene monoepoxide; and preferably vinyl substituted or alkyl substituted monoepoxide like 4-vinyl-1,2-epoxycyclohexane, 4-vinyl-1,2-epoxynorbornene, or limonene monoepoxide. The amount of cycloaliphatic mono- or diepoxide present in any suitable, including known, from 0 to about 40, preferably from about 1 to about 30, and most preferably from about 1 to about 20 weight percent based on the amount of the epoxidized cycloaliphatic ester.

If desired, minor amounts of glycidyl epoxides such as the diglycidyl ethers of Bisphenol-A, diglycidyl ethers of brominated Bisphenol-A, cresol-novolac epoxy resins, epoxy phenol novolac resins, diglycidyl ethers of 1,4-butanediol, and the like can be combined with the epoxidized cycloaliphatic esters.

Radiation-curable compositions may contain substituted and unsubstituted, linear or cyclic, vinyl ethers, such as acrolein dimer, acrolein tetramer, 2-methoxy tetrahydropyran, pyran, 2-methoxy dihydropyran, triethylene glycol divinyl ether, 1,4-cyclohexanedimethanol divinyl ether, butyl vinyl ether, and others.

Photoinitiators which may be used in the photocurable compositions include, among others, one or more: metal fluoroborate; complex of boron trifluoride, as described in U.S. Patent No. 3,379,653; bis-(perfluoroalkyl-sulfonyl)methane metal salt, as described in U.S. Patent No. 3,586,616; aryl diazonium compound as described in U.S. Patent No. 3,708,296; aromatic onium salt of Group VIa elements as described in U.S. Patent No. 4,058,400; aromatic onium salt of Group Va elements as described in U.S. Patent No. 4,069,055; dicarbonyl chelate of a Group IIIa-Va element as described in U.S. Patent No. 4,068,091; thiopyrylium salt as described in U.S. Patent No. 4,139,655; Group VIb element in an  $\text{MF}_6$ -anion where M is selected from phosphorous, antimony, and arsenic as described in U.S. Patent No. 4,161,478; arylsulfonium complex salt as described in U.S. Patent No. 4,231,951; aromatic iodonium complex salt and aromatic sulfonium complex salt, as described in U.S. Patent 4,256,828; bis(4-(diphenylsulfonio)-phenyl) sulfide-bis-hexafluorometallic salts such as the phosphate, arsenate, antimonate and the like as described Watt et al. in J. Polymer Sci.: Polymer Chem. Ed., 22, 1789 (1984); and preferably cationic photoinitiators including arylsulfonium complex salts, aromatic sulfonium or iodonium salts of halogen containing complex ions, and aromatic onium salts of Group II, V, and VI elements, such as FX-512 (3M Co.), UVR-6990 and UVR-6974 (Union Carbide Chemicals and Plastics Co. Inc.), UVE-1014 and UVE-1016 (General Electric Co.), KI-85 (Degussa), and SP-150 and SP-170 (Asahi Denka). The photocurable composition may also contain benzophenone or a derivative of benzophenone.

These photoinitiators generate both cations, which cause polymerization of the epoxidized cycloaliphatic esters, and free radicals, which can be used to polymerize acrylates, such that photocurable compositions may also contain acrylates. Suitable acrylates may be derived from ethylenically unsaturated monomers such as: esters of acrylic and methacrylic acid with monohydric and polyhydric compounds, such as methyl, ethyl, propyl, butyl, pentyl, hexyl, octyl, decyl, and the like acrylates and methacrylates; like neopentyl diacrylate, esterdiol diacrylates such as 2,2-dimethyl-3-hydroxypropyl-2,2-dimethyl-3-hydroxypropionate diacrylate, trimethylolpropane triacrylate, pentaerythritol di-, tri-, and tetraacrylate, hydroxyethyl acrylate, hydroxypropyl acrylate, caprolactone acrylates, ethoxylated acrylates, propoxylated acrylates, glycerol acrylates, triethylene glycol diacrylate, tetraethylene glycol diacrylate, ethoxyethyl acrylate, cyclohexyl acrylate, 2-phenoxyethyl acrylate, isobornyl acrylate, 1,3-butylene glycol diacrylate, 1,4-butanediol diacrylate, 1,6-hexanediol diacrylate, glycidyl acrylate, or the corresponding methacrylates; styrene; divinylbenzene; N-vinylpyrrolidone; and others. Oligomers or polymers which can be used in photopolymerizable compositions include, among others: poly(ethylene glycol) acrylates; caprolactone di-, tri-, and tetraacrylates; tripropylene glycol diacrylate; poly(propylene glycol) acrylates; ethoxylated or propoxylated Bisphenol A diacrylates; alkoxylated esterdiol diacrylates, such as ethoxylated or propoxylated 2,2-dimethyl-3-hydroxypropyl-2,2-dimethyl-3-hydroxypropionate diacrylates; acrylates of caprolactone reacted with esterdiols; ethoxylated or propoxylated trimethylolpropane triacrylate; ethoxylated or propoxylated pentaerythritol di-, tri-, or tetraacrylate; unsaturated polyesters containing ethylenic unsaturation from maleic, fumaric, citraconic, or other unsaturated dicarboxylic acids; urethane acrylates; epoxy acrylates; acrylated polybutadiene; acrylated linseed oil; acrylated soybean oil; and others. Photocurable compositions may also contain homolytic fragmentation-type, free radical-generating photoinitiator when acrylates are present.

Photopolymerization may be carried out by exposing the film or coating containing the unsaturated cycloaliphatic esters or derivatives to electromagnetic radiation which is rich in short-wave radiation. Particularly useful is radiation of about 200 to about 400 nanometers in wavelength. Illustrative of appropriate light sources are low pressure, medium pressure, and high pressure mercury vapor lamps, xenon and other flash-type lamps, fluorescent lights, lasers, electrodeless mercury lamps, and the like. Other sources of radiant energy such as electron beams, gamma radiation, X-rays, sunlight, and so on can also be used.

The photocurable compositions may contain, exclusive of photoinitiator, from about 25 to 100 percent of the cycloaliphatic epoxides, from 0 to about 60 percent of hydroxyl-containing compound, from 0 to about 75 percent of other cycloaliphatic or other epoxide, from 0 to about 60 percent vinyl ether, and from 0 to about 60 percent acrylate. The photocurable compositions may also contain other ingredients such as one or more surfactants, flow and leveling agents, fumed silicas, silicone oils and other slip agents, and other ingredients suitable for coatings.

The thermally-curable compositions can contain suitable, including known, catalysts such as sulfuric acid, hydrochloric acid, p-toluene sulfonic acid, methyl sulfonic acid, phosphoric acid and alkyl derivatives of phosphoric acid, maleic acid, trimellitic acid, triflic acid, salts of triflic acid such as the diethylammonium salt of triflic acid, the ammonium salt of triflic acid, the stannous salt of triflic acid, stannous octanoate, uranyl nitrate, zinc octanoate, and the like, including mixtures of these catalysts. The thermally-curable compositions may contain, exclusive of catalyst, from about 25 to 100 percent of the cycloaliphatic epoxides, from 0 to about 60 percent of a hydroxyl-containing compound, and from 0 to about 75 percent of other cycloaliphatic or other epoxide. The thermally-curable compositions may also contain other ingredients such as one or more surfactants, flow and leveling agents, fumed silicas, silicone oils and other slip agents, and other ingredients suitable for coatings.

The crosslinkable coating compositions can also contain other suitable ingredients like pigments, fillers, surfactants, flow and leveling agents, fumed silica, slip agents, and other additives useful in coating compositions, in suitable, including known, quantities. Selection of particular coating additives may follow established practice. In preparing the crosslinkable polymeric coating compositions, the ingredients can be mixed by any suitable means, including conventional procedures used in the production of paint, ink, adhesive, and sealant compositions. The coating compositions may be applied to a surface or substrate by any suitable, including conventional, means. Thermal curing can be conducted by heating at a suitable temperature generally from about 50 °C to about 275 °C, preferably from about 90 °C to about 200 °C, for a period of time sufficient to obtain a dry film. Generally this time will range from about one minute to about two hours. The components present in a particular crosslinkable polymeric coating composition will determine the temperature and time that will be required to obtain an adequate cure and a good coating film.

### Examples

The following examples present illustrative embodiments of this invention and are not intended to limit its scope. All of the parts, percentages and proportions referred to herein, including the claims, are by weight unless otherwise indicated. The following terms used in the examples have the following meanings:

Designation	Description
Photoinitiator I	Hexafluoroantimonate sulfonium salt, available as CYRACURE™ UVI-6974 from Union Carbide Chemicals and Plastics Company Inc.
Polyol I	Propylene oxide polyol with an average hydroxyl number of 112 and an average equivalent weight of 500, available as NIAX™ Polyol LHT-112 from Union Carbide Chemicals and Plastics Company Inc.
Polyol II	Trihydroxyl functional ε-caprolactone polyol with an average hydroxyl number of 312 and an average equivalent weight of 180, available as TONE™ -0305 from Union Carbide Chemicals and Plastics Company Inc.
Epoxide I	3,4-Epoxy cyclohexanemethyl 3,4-epoxycyclohexanecarboxylate, available as CYRACURE™ UVR-6110 from Union Carbide Chemicals and Plastics Company Inc.
Epoxide II	1-Vinyl 3,4-epoxycyclohexane.
Surfactant I	A silicone-alkylene oxide copolymer, available as SILWET™ L-7604 from Union Carbide Chemicals and Plastics Company Inc.

Measurements and test procedures used in the examples are as follows:

**Solvent Resistance (Double Acetone Rubs):** A measure of the resistance of the cured film to attack by acetone in which a film coating surface was rubbed with an acetone-soaked cloth back and forth with hand pressure. A rub back and forth over the film coating surface with the acetone-soaked cheesecloth was designated as one "double acetone rub".

**Pencil Hardness:** Pencil leads of increasing hardness values were forced against the film coating

surface in a precisely defined manner as described in ASTM D3363-74 until one pencil lead cut through the surface of the film coating. The surface hardness was considered as the hardest pencil grade which just failed to cut or mar the film coating surface. The pencil leads in order of softest to hardest were reported as follows: 6B, 5B, 4B, 3B, 2B, B, HB, F, H, 2H, 3H, 4H, 5H, 6H, 7H, 8H, and 9H.

5 Crosshatch Adhesion. A lattice pattern with ten cuts in each direction was made in the coating film to the substrate and pressure-sensitive adhesive tape (Scotch Brand 606) was applied over the scored/cut substrate and then quickly removed. The amount of coating remaining on the scored area is the "Percent Crosshatch Adhesion".

10 Gardner Impact Resistance. A measure of the ability of a cured film coating on a substrate to resist rupture from a falling weight. A Model IG-1120 Gardner Impact Tester equipped with an eight-pound dart was used to test film coatings cast and cured on steel panels. The dart was raised to a given height in inches and dropped onto either the coated side of the coated steel panel (direct or forward impact resistance) or the uncoated side of the coated steel panel (reverse impact resistance). The height-of-drop in inches times weight of dart (8 pounds), designated as inch-pounds, absorbed by the film without rupturing  
15 was recorded as the films direct or reverse impact resistance.

#### Examples 1-17: Cycloaddition

20 These examples describe reacting dienes with dienophilic (meth/eth)acrylates to produce unsaturated cycloaliphatic esters. Specific reactants, and amounts, are given in Table A, using the following procedures.

In Example 1, 100 ml of toluene solvent and 5 grams (0.037 mole) of anhydrous aluminum chloride catalyst were added to a three-necked, one-liter, round-bottom flask equipped with a water-cooled reflux condenser, thermometer, and addition funnel. The mixture was kept under a nitrogen blanket and the flask contents were stirred as the indicated amount of dienophilic ester given in Table A dissolved in 50 ml of  
25 toluene was added dropwise over a 30-minute period. This was followed by the dropwise addition of the indicated amount of diene given in Table A in 25 ml of toluene over a one-hour period. The reaction temperature was maintained at 30° C. by occasionally cooling the reaction flask with cold water. After an additional two hours of stirring, the reactor contents were poured, with stirring, onto a mixture of 200 ml of ice and 10 ml of concentrated hydrochloric acid. The organic layer was then separated from the water layer  
30 and washed successively with 100 ml portions of a 5% aqueous hydrochloric acid solution and water. The reaction was upscaled by a factor of 2.5 and carried out two additional times to produce a total of 1328 grams of crude product with the following composition: 1.4% diene, 69% toluene, and 29% product. The mixture was transferred to a three-liter, round-bottom flask equipped with a 20-tray Oldershaw column and a water-cooled automatic reflux head. Distillation under vacuum gave 381 grams of product (86% yield based  
35 on combining ratios of original reactants) with a 99% purity and a boiling point of 160° C. at 145 mm Hg. The low viscosity product had a pleasing fragrance, and was stored for future use.

In Examples 2-13, the procedures of Example 1 is followed except that the dienophiles and dienes indicated in Table A are used.

40 In Examples 14-15, 200 ml of toluene solvent and 10 grams (0.037 mole) of anhydrous aluminum chloride catalyst are added to a three-necked, two-liter, round-bottom flask equipped with a water-cooled reflux condenser, thermometer, and addition funnel. The mixture is kept under a nitrogen blanket, and the flask contents cooled to 0° C. and stirred as the indicated amount of dienophilic ester given in Table A dissolved in 100 ml of toluene is slowly added over a 45-minute period. This is followed by the slow  
45 addition of the indicated amount of diene given in Table A in 100 ml of toluene over a two-hour period. The reaction temperature is slowly increased to 25° C. by heating or cooling the reaction flask with cold water. After stirring for an additional four hours, the reactor contents are poured, with stirring, onto a mixture of 400 grams of ice and 20 ml of concentrated hydrochloric acid. The organic layer is then separated from the water layer and washed successively with 100 ml portions of a 5% aqueous hydrochloric acid solution and  
50 of water. The product layer is then stripped at 70° C. and aspirator vacuum to remove any low molecular weight volatiles and toluene.

In Example 16, the procedure of Example 2 is followed using the amounts of diene and dienophilic ester given in Table A except for the following variations. The mixture is instead cooled to -5° C. and the reaction temperature is slowly increased to 5° C. and held there for 30 minutes and then slowly increased to 10° C. and held there for 60 minutes. Temperature is maintained by occasionally cooling the reaction flask  
55 with cold water or heating with a heat gun if necessary. After stirring for an additional two hours, the product is isolated using the same procedure.

In Example 17, the procedure of Example 16 is followed using the amounts of diene and dienophilic ester given in Table A except for the following variations. The mixture is instead kept under a nitrogen

blanket at 25°C. and the reaction temperature held at 25°C. for 3 hours and then slowly increased to 40°C. and held there for 60 minutes. The same procedure is followed to produce and isolate product except that 600 grams of ice are used.

TABLE A. Cycloaddition Reactions

Ex.	Diene (gms, moles)	Dienophile (gms, moles)	n	x	R <sub>1</sub>	R <sub>2</sub>	R <sub>3</sub>	R <sub>4</sub>	R <sub>5</sub>	R <sub>6</sub>	R <sub>7</sub>	Cycloaliphatic Ester <sup>a</sup>			
												R <sub>8</sub>	R <sub>9</sub>	R <sub>10</sub>	
1	Isoprene(34,0.5)	Ethyl acrylate(44,0.44)	1	0	H	H	H,CH <sub>3</sub>	CH <sub>3</sub>	H	H	-	-	H	-C <sub>2</sub> H <sub>5</sub>	
2	1,3-Butadiene(81,1.5)	HEA <sup>d</sup> (116,1)	1	0	H	H	H	H	H	H	-	-	H	-CH <sub>2</sub> CH <sub>2</sub> OH	
3	1,3-Butadiene(81,1.5)	HPA <sup>e</sup> (130,1)	1	0	H	H	H	H	H	H	-	-	H	Note f	
4	1,3-Butadiene(81,1.5)	HEMA <sup>f</sup> (130,1)	1	0	H	H	H	H	H	H	-	-	CH <sub>3</sub>	-CH <sub>2</sub> CH <sub>2</sub> OH	
5	Isoprene(102,1.5)	HPMA <sup>g</sup> (144,1)	1	0	H	H	H,CH <sub>3</sub>	CH <sub>3</sub>	H	H	-	-	CH <sub>3</sub>	Note j	
6	1,3-Butadiene(270,5)	DPPA <sup>h</sup> (525,1)	5	0	H	H	H	H	H	H	-	-	H	Note k	
7	1,3-Butadiene(216,4)	DPPA <sup>h</sup> (525,1)	4	0	H	H	H	H	H	H	-	-	H	Note l	
8	1,3-Butadiene(162,3)	DPPA <sup>h</sup> (525,1)	3	0	H	H	H	H	H	H	-	-	H	Note m	
9	1,3-Butadiene(108,2)	DPPA <sup>h</sup> (525,1)	2	0	H	H	H	H	H	H	-	-	H	Note n	
10	1,3-Butadiene(54,1)	DPPA <sup>h</sup> (525,1)	1	0	H	H	H	H	H	H	-	-	H	Note o	
11	1,3-Butadiene(27,0.5)	DPPA <sup>h</sup> (525,1)	0.5	0	H	H	H	H	H	H	-	-	H	Note p	
12	2,4-Hexadiene(41,0.5)	Ethyl acrylate(44,0.44)	1	0	CH <sub>3</sub>	H	H	H	H	CH <sub>3</sub>	-	-	H	-C <sub>2</sub> H <sub>5</sub>	
13	1,3-Hexadiene(41,0.5)	Ethyl acrylate(44,0.44)	1	0	C <sub>2</sub> H <sub>5</sub>	H	H	H	H	H	-	-	H	-C <sub>2</sub> H <sub>5</sub>	
14	1,3-Butadiene(162,3)	TMPTA <sup>i</sup> (296,1)	3	0	H	H	H	H	H	H	-	-	H	Note q	
15	Isoprene(238,3.5)	TMPTA <sup>i</sup> (296,1)	3	0	H	H	H,CH <sub>3</sub>	CH <sub>3</sub>	H	H	-	-	H	Note q	
16	Cyclo-C <sub>5</sub> <sup>b</sup> (231,3.5)	TMPTA <sup>i</sup> (296,1)	1-3	0	-	H	H	H	H	-	H	H	H	Note r	
17	Cyclo-C <sub>5</sub> TMS <sup>c</sup> (484,3.5)	TMPTA <sup>i</sup> (296,1)	1-3	0	-	H	H	H	H	-	H	Si(CH <sub>3</sub> ) <sub>3</sub>	H	Note r	

## Notes for Table A:

a - Having a structure as shown in Figure 1 with the noted substituents.

b - Cyclo-C<sub>5</sub> = 1,3-cyclopentadiene

c - Cyclo-C<sub>5</sub>TMS = 1,3-cyclopentadiene-5-yltrimethylsilane

d - HEA = hydroxyethyl acrylate

e - HPA = hydroxypropyl acrylate

f - HEMA = hydroxyethyl methacrylate

g - HPMA = hydroxypropyl methacrylate

h - DPPA = dipentaerythritol pentacrylate

i - TMPTA = trimethylolpropane triacrylate

j - -CH<sub>2</sub>CH(CH<sub>3</sub>)OH

k - [(-CH<sub>2</sub>)<sub>aa</sub>(HOCH<sub>2</sub>)<sub>bb</sub>(CH<sub>2</sub>=CHCOOCH<sub>2</sub>)<sub>cc</sub>CCH<sub>2</sub>]<sub>2</sub>O  
wherein aa is 5, bb is 1 and cc is 0.

l - [(-CH<sub>2</sub>)<sub>aa</sub>(HOCH<sub>2</sub>)<sub>bb</sub>(CH<sub>2</sub>=CHCOOCH<sub>2</sub>)<sub>cc</sub>CCH<sub>2</sub>]<sub>2</sub>O  
wherein aa is 4, bb is 1 and cc is 1.

m - [(-CH<sub>2</sub>)<sub>aa</sub>(HOCH<sub>2</sub>)<sub>bb</sub>(CH<sub>2</sub>=CHCOOCH<sub>2</sub>)<sub>cc</sub>CCH<sub>2</sub>]<sub>2</sub>O  
wherein aa is 3, bb is 1 and cc is 2.

n - [(-CH<sub>2</sub>)<sub>aa</sub>(HOCH<sub>2</sub>)<sub>bb</sub>(CH<sub>2</sub>=CHCOOCH<sub>2</sub>)<sub>cc</sub>CCH<sub>2</sub>]<sub>2</sub>O  
wherein aa is 2, bb is 1 and cc is 3.

o - [(-CH<sub>2</sub>)<sub>aa</sub>(HOCH<sub>2</sub>)<sub>bb</sub>(CH<sub>2</sub>=CHCOOCH<sub>2</sub>)<sub>cc</sub>CCH<sub>2</sub>]<sub>2</sub>O  
wherein aa is 1, bb is 1 and cc is 4.

p - [(-CH<sub>2</sub>)<sub>aa</sub>(HOCH<sub>2</sub>)<sub>bb</sub>(CH<sub>2</sub>=CHCOOCH<sub>2</sub>)<sub>cc</sub>CCH<sub>2</sub>]<sub>2</sub>O  
wherein aa is 0.5, bb is 1 and cc is 4.5.

q - (-CH<sub>2</sub>)<sub>3</sub>CCH<sub>2</sub>CH<sub>3</sub>

r - (-CH<sub>2</sub>)<sub>n</sub>C(-CH<sub>2</sub>OCCH=CH<sub>2</sub>)<sub>3-n</sub>  

$$\begin{array}{c} | \\ \text{CH}_2\text{CH}_3 \quad \text{O} \end{array}$$

#### Examples 18-40: Epoxidation and Other Derivatizations

These examples describe various reactions of unsaturated cycloaliphatic esters including epoxidation, hydrogenation, (trans)esterification, and other derivatizations. Specific reactants and amounts are given in Table B using the following procedures. In Examples 18, 19, 37, 38 and 47, unsaturated cycloaliphatic esters were epoxidized as shown in Equation VII. In Examples 20-35, 43-46, 55-57, 63-65, 73 and 75, unsaturated cycloaliphatic esters are epoxidized, as shown in Equation VII. In Examples 48-50 and 57-59, unsaturated cycloaliphatic esters are hydrogenated, as shown in Equation VIII. In Examples 69-71, unsaturated cycloaliphatic esters having active hydrogen are further esterified, as shown in Equation IX. In Example 74, unsaturated cycloaliphatic esters are transesterified, as shown in Equation VI. Other derivatizations are



shown in Examples 39-42, 51-53, 60-62 and 66-68.

In Example 18, 50 ml of methylene chloride solvent and 36.6 grams (0.200 mole) of the Example 1 product were added to a three-necked, two-liter, round-bottom flask equipped with a water-cooled reflux condenser, thermometer, and dropping funnel. The compounds were well mixed. A solution containing 42 grams (0.243 mole) of meta-chloroperoxybenzoic acid epoxidizing agent and 800 ml of methylene chloride was prepared in a glass container. This solution was added to the reaction flask through the dropping funnel over a 65-minute period with stirring. The temperature was initially 25°C. and is 29°C. at the end of the addition. The solution temperature dropped to 25°C. over a 90-minute period. Then, the system was warmed to 40°C. and maintained at this temperature for 2.5 hour. Heating was discontinued and the reaction mass was allowed to stand under ambient conditions overnight (about 17 hours). The cured product was cooled by placing the reaction flask in an ice-water bath, and then filtered through a Buchner funnel to remove the meta-chlorobenzoic acid that formed during the epoxidation. The filtered methylene chloride containing the epoxide was then placed in a two-liter separatory funnel and washed with saturated sodium bicarbonate solution, with the pH of the organic layer and the neutralized bicarbonate analyzed until the organic layer was neutral. Then, the organic layer was washed with 50 ml of saturated, aqueous sodium chloride solution. The organic layer was dried over anhydrous magnesium sulfate for 45 minutes and filtered. The methylene chloride solvent was removed by simple distillation until the volume was about 100 ml. Chromatographic analysis of the product indicated that it contained 18.42% methylene chloride, 35% starting unsaturated ethyl ester, and 45.81% of a mixture of the corresponding epoxidized compounds, such as shown by Equation VII. The product was then vacuum distilled from a 200 ml round-bottom flask equipped with a 37 cm column packed with 1/8-inch glass helices. The fraction boiling at 126-137°C. and 4 mm Hg pressure was collected as the desired epoxide mixture. Chromatographic analysis indicated the product was about 97% of the mixed epoxide isomer and 2.67% of the starting unsaturated ethyl ester.

In Example 19, an unsaturated cycloaliphatic ester was initially made using a 500 ml, round-bottom, glass reaction flask equipped with a stirrer, a temperature measuring device, and a 20-tray Oldershaw column and decanting head which was charged with 53 grams (0.42 moles) of 3-cyclohexene-1-carboxylic acid, 116 grams (2.5 moles) of absolute ethanol solvent, 100 ml of cyclohexane azeotroping solvent, and 0.30 grams of sodium hydrogen sulfate esterification catalyst. The stirred reaction mass was brought to reflux and the upper layer of the water/ethanol/cyclohexane azeotropic mixture was returned to the column. After a reaction time of 12 hours, chromatographic analysis indicated that the 3-cyclohexene-1-carboxylic acid had been quantitatively converted into the corresponding ethyl ester. The excess ethanol was then evaporated under a vacuum of 100 mm Hg. The reaction mixture was diluted with cyclohexane, 50 ml, and successively washed with 50 ml portions of a saturated sodium bicarbonate solution and water. The organic layer was dried over anhydrous magnesium sulfate and transferred to a 200 ml flask and distilled under vacuum through a 37 cm column packed with one-eighth-inch glass helices to yield 47 grams (73% yield) of the desired ethyl ester (b.p. 116-118°C. at 100 mm Hg) with a purity exceeding 99%.

In Example 19, epoxidation was conducted using a 3-neck, 200 ml, glass round-bottom flask was equipped with a water-cooled reflux condenser, thermometer, and addition funnel and charged with 45 grams (0.29 mole) of the product of the designated examples. While stirring, 90 grams (0.29 mole) of a 25% solution of peracetic acid epoxidizing agent in ethyl acetate solvent was added dropwise to the reaction flask over a one-hour period. The reaction temperature was maintained below 80°C. by cooling the flask with cold water. After all the peracetic acid solution had been added, the reaction mass was stirred for an additional 30 minutes and then washed twice with 60 ml portions of water followed by two washes, 60 ml portions, with an aqueous saturated sodium bicarbonate solution. The organic layer was dried over anhydrous magnesium sulfate, filtered, and transferred to a 200 ml round-bottom flask. The product was distilled under vacuum through a 37 cm column packed with one-eighth inch glass helices to give 28 grams (57% yield) of ethyl 3,4-epoxycyclohexanecarboxylate, b.p. 87-88°C. at 1.0 mm Hg, purity exceeding 99% was provided.

In Examples 20-35, the epoxidation procedure of Example 19 is followed except that the unsaturated esters identified in Table B are used.

Example 36 describes the preparation of 2-propoxyethyl 3-cyclohexenecarboxylate. A 500 ml, round-bottom flask equipped with a 10-tray Oldershaw column and decanting head, thermometer, and a connecting tube with stopcock was charged with 126 grams (1.0 mole) of 3-cyclohexene-1-carboxylic acid, 99 grams (0.95 mole) of 2-propoxyethanol, and 100 ml of toluene. The reaction mass was stirred and brought to reflux. Then 1.1 grams of titanium tetrabutoxide was added via the connecting tube. The temperature of the reaction mass was slowly increased to 200°C. and held at this temperature while removing water and solvent overhead. After a reaction time of 8 hours, the product was cooled and washed with 50 ml of dilute phosphoric acid (10 weight percent in water) and filtered to remove the titanium salts.

The organic layer was separated from the water layer, and the organic layer was successively washed with 50 ml portions of a saturated sodium bicarbonate solution and with water. Distillation of the crude product under vacuum through a 12-cm Vigreux column gave 152 grams (75% yield) of 2-propoxyethyl 3-cyclohexenecarboxylate (b.p. 88-90° C. at 2.0 mm Hg) with a chromatographically analyzed purity exceeding 98%.

In Example 37, butyl bis(3-cyclohexenecarboxylate) was made using the procedure described in Example 36 by reacting 124 grams (0.98 mole) of 3-cyclohexene-1-carboxylic acid with 43.2 grams (0.48 mole) of 1,4-butanediol. Distillation of the crude product resulted in a yield of 121 grams (83% yield) of butyl bis(3-cyclohexene carboxylate) (b.p. 223-225° C. at 2.0 mm Hg) with an chromatographically analyzed purity exceeding 99%. The product was epoxidized by the same procedure used in Example 19. Chromatographic analysis indicated the corresponding diepoxide was obtained.

In Example 38 trimethylolpropane tris(3-cyclohexenecarboxylate) and the corresponding triepoxide were made. In the same manner as described in Example 36, 124 grams (0.98 mole) of 3-cyclohexenecarboxylic acid was reacted with 43 grams (0.32 mole) of trimethylolpropane (2-ethyl-2-hydroxymethyl)-1,3-propanediol. Purification of the crude product yielded 139 grams (95% yield) of trimethylolpropane tris(3-cyclohexenecarboxylate) in the form of a pale yellow, viscous liquid. A portion of the product was epoxidized by the same procedure as that used in Example 19, and chromatographic analysis indicated that the product was trimethylolpropane tris(3,4-cyclohexane carboxylate).

In Examples 39-42, 0.20 mole of the ester given in Table B are placed in a glass reaction flask equipped with a stirrer, thermometer, nitrogen inlet and outlet, and an addition tube. The ester is heated to 60° C. and held at this temperature for one hour with stirring and while sparging with dry nitrogen. Then, 0.1 mole of the designated diisocyanate is slowly added with stirring over a 30-minute period at the temperature of 60° C. The temperature is controlled between 55° C. and 65° C. by use of either an ice/water bath or a heating gun for a time period sufficient for the isocyanato groups to react with the hydroxyl groups on the ester. Completion of reaction is either determined by infrared analysis for residual isocyanate or by titration for free isocyanate.

In Examples 43-46, the products of Examples 39-42 are epoxidized by the same procedure as described in Example 31.

In Example 47, 140 grams (1.1 moles) of 3-cyclohexene-1-carboxylic acid were reacted with 46 grams (0.18 mole) of dipentaerythritol using the procedures as described in Example 36. Evaporation of the toluene and purification of the crude product gave 150 grams (92% yield) of the corresponding hexafunctional unsaturated product which was epoxidized following the procedure described in Example 19.

In Examples 48-71, unsaturated cyclohexene esters of Examples 2-4 are each alone or in combination reacted in either neat form or in a solvent diluted form to make derivative compounds. In Examples 48-50, the esters are hydrogenated to form saturated, monohydroxyl-functional compounds as shown in Equation VIII. In Examples 51-53, the esters are reacted with monoisocyanate, such as methyl isocyanate, butyl isocyanate or phenyl isocyanate, to form the corresponding unsaturated cyclohexene urethanes. The urethanes of Examples 51-53 are epoxidized in Examples 54-56 following the procedure described in Example 31 to form the corresponding cycloaliphatic epoxides and are hydrogenated in Examples 57-59 to form the corresponding saturated urethanes. In Examples 60-62, one mole of each ester is reacted with 1 to 50 moles of ethylene oxide, propylene oxide, butylene oxide, epichlorohydrin, or  $\epsilon$ -caprolactone each alone or in combination to form the corresponding unsaturated cyclohexene adducts. In Examples 63-65, these adducts are epoxidized following the procedure described in Example 31. In Examples 66-68, the epoxides of Example 63-65 are reacted with monoisocyanates using the procedure as described in Example 51-53. In Examples 69-71, two moles of the adducts of Examples 60-62 are reacted with an anhydride and/or a carboxylic acid or polycarboxylic acid, such as maleic anhydride or acid, phthalic anhydride or acid, isophthalic anhydride, trimellitic anhydride, hexahydrophthalic anhydride, or other anhydride or carboxylic acid to form bis(cyclohexene) compounds that are linked by a maleate linkage in the case of maleic anhydride, or other linkage indicative of the specific acid or anhydride used.

In Example 72, transesterification is described for preparing other ester compounds, as shown in Equation VI. One mole of the product of Example 1, 2 moles of butanol (1.0 mole excess), and 0.5 weight percent tetrabutyltitanate transesterification catalyst, are placed in a suitable reaction flask equipped with a thermometer, stirrer, and condenser. The contents of the flask are heated to 160 to 200° C. and held at these temperatures for 2 to 20 hours. Ethanol and excess butanol are removed by distillation, and the corresponding butyl ester product is recovered by further distillation.

In Example 73, the butyl ester of Example 72 is epoxidized by the process described in Example 31.

Example 74, describes preparation of an octafunctional unsaturated compound. In the same manner as described in Example 36, 127 grams (1.0 mole) of 3-cyclohexene-1-carboxylic acid is reacted with 39.8

grams (0.125 mole) of tripentaerythritol. The resulting reaction product, tripentaerythritol octa-(cyclohexenecarboxylate) is formed as a viscous compound.

In Example 75 the product of Example 74 epoxidized by the procedure described in Example 31.

Table B. Derivatisations

Ex. Product	Derivatizing Agent	Starting Ester	Ex. Product	Derivatizing Agent	Starting Ester
18 Epoxide	m-CPBA	Ex. 1	47 Epoxide	Peracetic acid	(in Ex. 47)
19 Epoxide	Peracetic acid (in Ex. 19)	48 Hydrot.	Hydrogen	Ex. 2	
20 Epoxide	Peracetic acid	Ex. 2	49 Hydrot.	Hydrogen	Ex. 3
21 Epoxide	Peracetic acid	Ex. 3	50 Hydrot.	Hydrogen	Ex. 4
22 Epoxide	Peracetic acid	Ex. 4	51 Urethane	Isocyanate	Ex. 2
23 Epoxide	Peracetic acid	Ex. 5	52 Urethane	Isocyanate	Ex. 3
24 Epoxide	Peracetic acid	Ex. 6	53 Urethane	Isocyanate	Ex. 4
25 Epoxide	Peracetic acid	Ex. 7	54 Epoxide	Peracetic Acid	Ex. 51
26 Epoxide	Peracetic acid	Ex. 8	55 Epoxide	Peracetic Acid	Ex. 52
27 Epoxide	Peracetic acid	Ex. 9	56 Epoxide	Peracetic Acid	Ex. 53
28 Epoxide	Peracetic acid	Ex. 10	57 Hydrot.	Hydrogen	Ex. 51
29 Epoxide	Peracetic acid	Ex. 11	58 Hydrot.	Hydrogen	Ex. 52
30 Epoxide	Perpropionic acid	Ex. 12	59 Hydrot.	Hydrogen	Ex. 53
31 Epoxide	Peracetic acid	Ex. 13	60 Ether	Epoxides	Ex. 2
32 Epoxide	Peracetic acid	Ex. 14	61 Ether	Epoxides	Ex. 3
33 Epoxide	Peracetic acid	Ex. 15	62 Ether	Epoxides	Ex. 4
34 Epoxide	Peracetic acid	Ex. 16	63 Epoxide	Peracetic acid	Ex. 60
35 Epoxide	Peracetic acid	Ex. 17	64 Epoxide	Peracetic acid	Ex. 61
36 Ester	-	(in Ex. 36)	65 Epoxide	Peracetic acid	Ex. 62
37 Epoxide	Peracetic acid (in Ex. 37)	66 Urethane	Isocyanate	Ex. 63	
38 Epoxide	Peracetic acid (in Ex. 38)	67 Urethane	Isocyanate	Ex. 64	
39 Urethane IPDI <sup>b</sup>		Ex. 2	68 Urethane	Isocyanate	Ex. 65

Table B. Derivatizations (continued)

Ex. Product	Derivatizing Agent	Starting Ester		Ex. Product	Derivatizing Agent	Starting Ester
		Ex. 3	Ex. 4			Ex. 60
140 Urethane TDI <sup>c</sup>		69 (Di)ester Acid/anhydride	70 (Di)ester Acid/anhydride			Ex. 61
41 Urethane DPMDI <sup>d</sup>		Ex. 4	71 (Di)ester Acid/anhydride			Ex. 62
42 Urethane DCHMDI <sup>e</sup>		Ex. 5	72 Transester Butanol			Ex. 1
43 Epoxide Peracetic acid		Ex. 39	73 Epoxide Peracetic acid			Ex. 72
44 Epoxide Peracetic acid		Ex. 40	74 Ester			(in Ex. 74)
45 Epoxide Peracetic acid		Ex. 41	75 Epoxide Peracetic acid			Ex. 74
46 Epoxide Peracetic acid		Ex. 42				

Notes for Table B

- a - m-CPBA = meta-chloroperoxybenzoic acid  
 b - IPDI = isophorone diisocyanate  
 c - TDI = toluene diisocyanate  
 d - DPMDI = 4,4'-diphenylmethane diisocyanate  
 e - DCHMDI = 4,4'-dicyclohexanemethyl diisocyanate

Examples 76-92: Coatings

These examples describe coating compositions containing epoxides of unsaturated cycloaliphatic esters produced in previous examples. In Examples 76-80, the ingredients listed in Table C were placed in amber-colored, glass bottles, mixed well, and coated onto Bonderite 37 steel panels with a No. 20 wire-wound rod. The coatings were cured by passing them under a 300 watt/inch ultraviolet light source (Fusion Systems Type A lamp) at 10 feet per minute. One panel coated with each mixture was exposed to the ultraviolet light radiation and allowed to stand for 24 hour before testing. A second panel coated in the same manner was given a thermal post cure of 10 minutes at 100° C. after ultraviolet light exposure. The coatings were tested using the previously described test procedures with the results listed in Table D. After either ultraviolet light exposure alone or ultraviolet light exposure plus thermal post cure, the Example 76 coating remained tacky indicating that as would be expected, it probably had a glass transition temperature below room temperature. These examples demonstrate that the product of Example 18 was a reactive diluent for

ultraviolet light curable coatings and other systems. When used in such coatings, it yielded properties that were comparable to or better than (impact resistance) those that were obtained with a known reactive diluent, 1-vinyl-3,4-epoxycyclohexane. Since the Example 18 product had a residual tackiness, it has a potential for use in certain adhesives.

In Examples 81-83, the ingredients listed in Table C were placed in amber bottles, well mixed, and coated onto Bonderite 37 steel panels with a No. 20 wire-wound rod. The coatings were cured with a 300 watt per inch ultraviolet light source (Fusion Systems Type A lamp) at 10 feet per minute. One set of coated panels was only exposed to the ultraviolet light source and then allowed to stand at room temperature for 24 hours before testing. A second set of coated panels was given a 10 minute, 100 °C. thermal post cure after ultraviolet light exposure and then tested. Film thickness was 0.0008 inches (0.8 mil) for all samples tested.

In Examples 84-87, the ingredients listed in Table C were placed in glass bottles, well mixed, and coated onto Bonderite 37 steel panels with a No. 20 wire-wound rod. The coatings were cured with a 300 watt-per-inch ultraviolet light source, Fusion Systems Type V ultraviolet lamp) at 10 feet per minute. The resultant coatings had excellent adhesion and solvent resistance, were quite hard, and displayed good impact resistance when tested directly on the surface of the coating except for the coating of Example 86 which displayed a combination of high hardness and toughness with excellent impact resistance when tested both directly on the surface of the coated substrate and from the reverse side of the coated substrate.

In Examples 88-92, the ingredients listed in Table C were placed in glass bottles, well mixed, and coated onto Bonderite 37 steel panels with a No. 20 wire-wound rod. The coatings were cured with a 300 watt-per-inch ultraviolet light source, Fusion Systems Type V ultraviolet lamp) at 10 feet per minute. The resultant coatings were hard and had excellent adhesion and solvent resistance, and displayed good impact resistance when tested directly on the surface of the coating. The coating of Example 92 was not tested for impact resistance.

Table C

Coating Formulations						
Ex.	Cycloaliphatic Epoxide		Epoxide	Polyol	Photoinitiator	Surfactant
	Ex.	Amount				
76	18	4.83	---	---	0.15(I)	0.02(I)
77	18	1.45	8.20(I)	---	0.30(I)	0.05(I)
78	18	1.45	6.20(I)	2.00(I)	0.30(I)	0.05(I)
79	--	---	8.20(I)	---	0.30(I)	0.05(I)
			1.45(II)			
80	--	---	6.20(I)	2.00(I)	0.30(I)	0.05(I)
			1.45(II)			
81	19	1.83	---	---	0.15(I)	0.02(I)
82	19	1.45	8.20(I)	---	0.30(I)	0.05(I)
83	19	1.45	6.20(I)	2.00(I)	0.30(I)	0.05(I)
84	37	2.93	---	---	0.118(I)	---
85	37	2.75	1.96(I)	---	0.118(I)	---
86	37	2.93	---	0.70(II)	0.07(I)	0.02(I)
87	37	2.75	1.96(I)	0.60(II)	0.11(I)	0.15(I)
88	38	3.03	---	---	0.033(I)	0.025(I)
89	38	2.05	1.01(I)	---	0.028(I)	0.022(I)
90	38	2.04	0.54(I)	---	0.08(I)	0.024(I)
			0.50(II)			
91	38	0.695	0.042(II)	0.072(II)	0.008(I)	0.006(I)
92	38	0.286	0.076(I)	---	0.011(I)	0.003(I)
	19	0.217	0.070(II)			

Table D

Coating Ex.	Double Acetone Rubs	Cross-hatch Adhesion	Pencil Hardness	Gardner Impact Resistance	
				Direct	Reverse
76 <sup>b</sup>	---	---	--	--	--
77	100	92(75)	2H	15(25)	<5
78	93(100)	100	2H	75(100)	15(20)
79	100	92(98)	2H	25	<5
80	100	100	2H	75(50)	25(50)
81 <sup>c</sup>	--(3)	--(90)	--(6B)	--(<5)	--(25)
82	>100	98(80)	2H	<5	25
83	>100	100	F	5(<5)	50
84	>100	100	3H	75	<5
85	>100	100	3H	50	<5
86	>100	100	2H	275	250
87	>100	100	3H	75	<5
88	>100	100	H	25	<5
89	>100	100	4H	50	<5
90	>100	100	4H	50	<5
91	>100	100	H	50	<5
92	45	100	3H		

a - Different values obtained after thermal post curing are in parentheses.

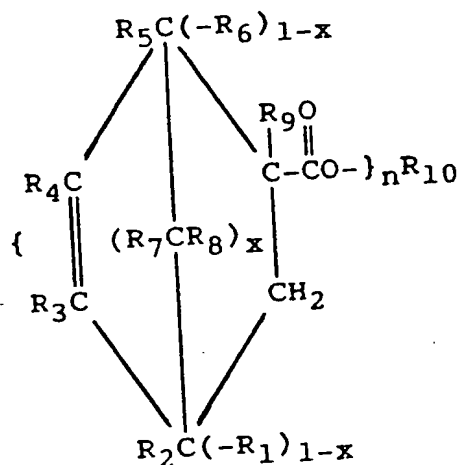
b - Sample is tacky and no properties were determined after ambient cure.

c - Low values after thermal cure reflect the uncrosslinked nature of this coating. Low pencil hardness of baked coating is evidence that the glass transition temperature was below the test temperature.

Although the invention has been illustrated by certain of the preceding examples, it is not to be construed as being limited thereby; but rather, the invention encompasses the generic area as hereinbefore disclosed. Various modifications and embodiments can be made without departing from the spirit and scope thereof.

### Claims

1. A process for producing an unsaturated cycloaliphatic ester represented by the formula:



wherein:

n is at least 1;

R<sub>1-8</sub> are each, independently, hydrogen, C<sub>1-10</sub> hydrocarbyl with or without halo substitution, halo, cyano, or silyl;

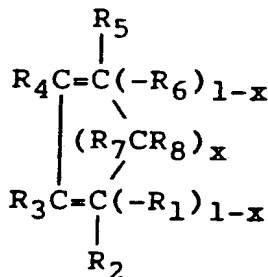
R<sub>9</sub> is hydrogen, methyl, or ethyl;

R<sub>10</sub> is hydrocarbyl or oxyhydrocarbyl, provided R<sub>10</sub> has at least 5 carbon atoms or is oxyhydrocarbyl or is -CH=CH<sub>2</sub> when n is 1; and

x is 0 or 1;

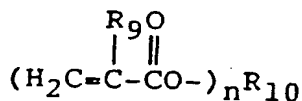
which comprises reacting:

(1) a diene represented by the formula:



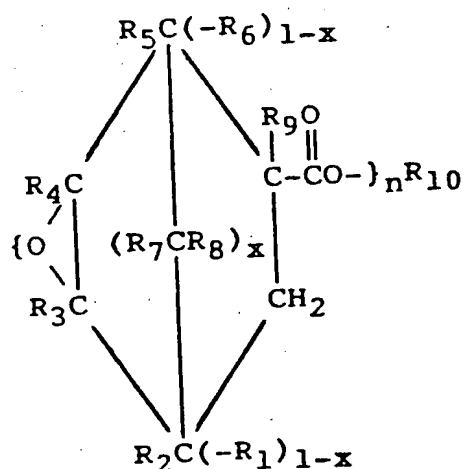
wherein R<sub>1-8</sub> and x are as defined previously, with:

(2) a dienophilic (meth/eth)acrylate represented by the formula:



wherein R<sub>9-10</sub> and n are as defined previously.

2. The process of Claim 1 wherein the diene is 1,3-butadiene, isoprene, 1-3-dicyclopentadiene, or 1-3-cyclopentadien-5-yltrimethylsilane.
3. The process of Claim 1 wherein the dienophilic (meth/eth)acrylate is trimethylolpropane triacrylate, hydroxyethyl acrylate, hydroxyethyl (meth)acrylate, hydroxypropyl acrylate, hydroxypropyl (meth)acrylate, dipentaerythritolpentaacrylate, pentaerythritol, dipentaerythritol, tripentaerythritol, or a polymer or oligomer containing free hydroxyl groups.
4. An unsaturated cycloaliphatic ester represented by the formula:



wherein:

n is at least 1;

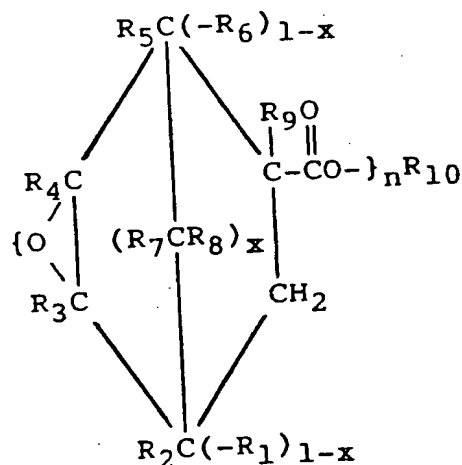
R<sub>1-8</sub> are each, independently, hydrogen, C<sub>1-10</sub> hydrocarbyl with or without halo substitution, halo, cyano; or silyl;

R<sub>9</sub> is hydrogen, methyl, or ethyl;

R<sub>10</sub> is hydrocarbyl or oxyhydrocarbyl, provided R<sub>10</sub> has at least 5 carbon atoms or is oxyhydrocar-

byl or is -CH=CH<sub>2</sub> when n is 1; and  
x is 0 or 1;

5. A cycloaliphatic epoxide represented by the formula:



wherein:

n is at least 1;

R<sub>1-8</sub> are each, independently, hydrogen, C<sub>1-10</sub> hydrocarbyl with or without halo substitution, halo, cyano, or silyl;

R<sub>9</sub> is hydrogen, methyl, or ethyl;

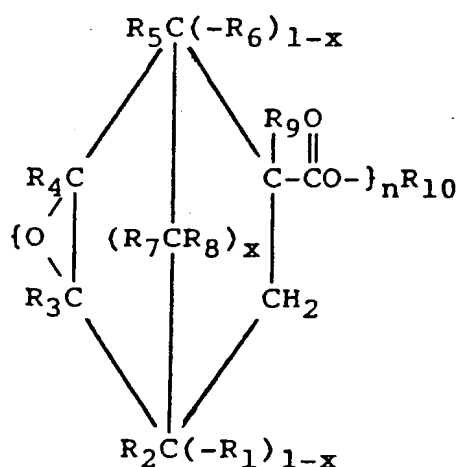
R<sub>10</sub> is hydrocarbyl or oxyhydrocarbyl, provided R<sub>10</sub> has at least 5 carbon atoms or is oxyhydrocar-

byl or is -CH=CH<sub>2</sub> when n is 1; and

x is 0 or 1;

6. A saturated cycloaliphatic ester represented by the formula:





wherein:

- n is at least 1;
- $R_1$ - $R_8$  are each, independently, hydrogen,  $C_{1-10}$  hydrocarbyl with or without halo substitution, halo, cyano, or silyl;
- $R_9$  is hydrogen, methyl, or ethyl;
- $R_{10}$  is hydrocarbyl or oxyhydrocarbyl, provided  $R_{10}$  has at least 5 carbon atoms or is oxyhydrocarbyl or is  $-CH=CH_2$  when n is 1; and
- x is 0 or 1;

7. An unsaturated cycloaliphatic ester which is the transesterification product of an unsaturated cycloaliphatic ester of Claim 4 reacted with an active hydrogen compound.
8. A cycloaliphatic epoxide which is the transesterification product of a cycloaliphatic epoxide of Claim 5 reacted with an active hydrogen compound.
9. A saturated cycloaliphatic ester which is the transesterification product of a saturated cycloaliphatic ester of Claim 6 reacted with an active hydrogen compound.

(19)



Europäisches Patentamt  
European Patent Office  
Office européen des brevets



(11) Publication number:

**0 520 419 A3**

(12)

**EUROPEAN PATENT APPLICATION**(21) Application number: **92110661.3**(51) Int. Cl.<sup>5</sup>: **C07C 67/347, C07D 303/40**(22) Date of filing: **25.06.92**(30) Priority: **26.06.91 US 721799**(43) Date of publication of application:  
**30.12.92 Bulletin 92/53**(84) Designated Contracting States:  
**AT BE CH DE DK ES FR GB GR IT LI LU MC  
NL PT SE**(88) Date of deferred publication of the search report:  
**16.06.93 Bulletin 93/24**(71) Applicant: **UNION CARBIDE CHEMICALS &  
PLASTICS TECHNOLOGY CORPORATION**  
**39 Old Ridgebury Road**  
**Danbury, Connecticut 06817-0001(US)**(72) Inventor: **Koleske, Joseph Victor**  
**1513 Brentwood Road**  
**Charleston 25314, West Virginia(US)**  
Inventor: **Argyropoulos, John Nicholas**  
**35 Michael Street**  
**Scott Depot 25560, West Virginia(US)**  
Inventor: **Smith, Oliver Wendell**  
**850 Walters Road**  
**Charleston 25314, West Virginia(US)**(74) Representative: **von Hellfeld, Axel, Dr.**  
**Dipl.-Phys.**  
**Wuesthoff & Wuesthoff Patent- und**  
**Rechtsanwälte Schweigerstrasse 2**  
**W-8000 München 90 (DE)**(54) **Production of unsaturated cycloaliphatic esters and derivatives thereof.**

(57) Unsaturated cycloaliphatic esters like higher hydrocarbyl, functionally substituted or polyunsaturated cyclohex-3-ene carboxylates, made directly by cycloaddition of dienes with dienophilic (meth/eth)acrylates, and their derivatives like epoxides and urethanes, provide useful thermal and radiation curable coatings, inks, sealants, adhesives, solvents, acid scavengers, and intermediates for other uses.

EP 0 520 419 A3



European Patent  
Office

# PARTIAL EUROPEAN SEARCH REPORT

which under Rule 45 of the European Patent Convention  
shall be considered, for the purposes of subsequent  
proceedings, as the European search report

Application Number

EP 92 11 0661  
Page 1

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int. Cl. 5)
X	GB-A-1 076 304 (CHISSO CORPORATION) ---	1	C07C67/347 C07D303/40
X	US-A-3 632 566 (THE LUBRIZOL CORPORATION) * complete document, in particular example 8 ; example 11 * ---	1,4	
X	CHEMICAL ABSTRACTS, vol. 80, 1974, Columbus, Ohio, US; abstract no. 14645c, FUKUMI,HIROKAZU ET AL. '4-Aminomethylcyclohexanecarboxylic acid esters' page 408 ; * abstract * * RN 50738-64-2 * & JP-A-73 052 747 (ASAHI CHEMICAL INDUSTRY CO., LTD.) ---	1,4	
D,X	US-A-2 794 812 (UNION CARBIDE AND CARBON CORPORATION) * example 3 * --- -/-	5	TECHNICAL FIELDS SEARCHED (Int. Cl. 5)  C07C C07D
INCOMPLETE SEARCH			
<p>The Search Division considers that the present European patent application does not comply with the provisions of the European Patent Convention to such an extent that it is not possible to carry out a meaningful search into the state of the art on the basis of some of the claims</p> <p>Claims searched completely : Claims searched incompletely : Claims not searched : Reason for the limitation of the search:</p> <p>see sheet C</p>			
Place of search THE HAGUE		Date of completion of the search 16 FEBRUARY 1993	Examiner VAN BIJLEN H.
CATEGORY OF CITED DOCUMENTS		I : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons ----- & : member of the same patent family, corresponding document	
X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document			



European Patent  
Office

# PARTIAL EUROPEAN SEARCH REPORT

Application Number

EP 92 11 0661

Page 2

DOCUMENTS CONSIDERED TO BE RELEVANT			CLASSIFICATION OF THE APPLICATION (Int. Cl. 5)
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	
X	DE-A-1 519 618 (BASF) * complete document * ---	6	
D,X	US-A-2 745 847 (UNION CARBIDE AND CARBON CORPORATION) * complete document * ---	4,5	
X	GB-A-822 980 (UNION CARBIDE CORPORATION) * example 1, 3 * ---	4,5	
X	GB-A-896 971 (UNION CARBIDE CORPORATION) * complete document * ---	4,5	
X	US-A-3 112 294 (SHELL OIL COMPANY) * example 4, 5 * ---	5	
X	FR-A-1 341 754 (BASF) * complete document * ---	4	TECHNICAL FIELDS SEARCHED (Int. Cl. 5)
X	FR-A-1 109 935 (UNION CARBIDE AND CARBON CORPORATION) * example 12, 13 * ---	4,5	
Y	US-A-4 310 681 (INTERNATIONAL FLAVORS & FRAGRANCES INC.) * complete document * ---	1	
Y	EP-A-0 037 644 (INTERNATIONAL FLAVORS & FRAGRANCES *) * complete document * ---	1	
Y	US-A-3 725 343 (E.I. DU PONT DE NEMOURS AND COMPANY) * example 7A * ---	1	
P,X	EP-A-0 479 166 (UNION CARBIDE CHEMICALS AND PLASTICS COMPANY, INC.) * complete document * -----	5	

## INCOMPLETE SEARCH

Claims searched completely : 1-3  
 Claims searched incompletely : 4-9

Concerning the scope of claims 4-6, the number of already known compounds (several of them are present in documents cited by the applicant himself) falling within the scope of these claims is so large that an exhaustive search for these claims is economically not justified (See Guidelines B III 2.1).

An on-line search for a stripped-down version of the generic formula's of claims 4-6 (i.e. without the -CR<sub>7</sub>R<sub>8</sub>-bridge and with R<sub>1</sub>-6 each hydrogen) revealed numerous compounds, part of which are listed here by their registry number:

138826-48-9	91976-21-5	62266-67-5
92319-47-6	103955-80-2	62266-66-4
118118-52-8	93157-93-8	62266-65-3
118118-51-7	92370-93-9	62266-62-0
118118-50-6	98684-59-4	62266-61-9
118117-85-4	94107-45-6	62266-59-5
118117-84-3	92773-34-7	62266-58-4
118117-83-2	90125-48-7	60787-70-4
114226-16-3	89611-20-1	53001-63-1
112406-93-6	77056-06-5	53001-59-5
93807-43-3	64465-50-5	53001-57-3

## OBSCURITIES

The structural formula of claims 4 and 6 do not correspond with an unsaturated, respectively with a saturated cycloaliphatic ester.

For the purpose of the search, based a.o. upon the description,

- for claim 4 (unsaturated esters), the structural formula of claim 1 was considered to be the correct and obvious formula;

- for claim 6 (saturated esters), the structural formula of claim 1 but with a single bond between

$$\begin{array}{c} \text{C} - \text{C} - \\ | \quad | \\ \text{R}_4 \quad \text{R}_3 \end{array}$$

was considered to be the correct and obvious structural formula.